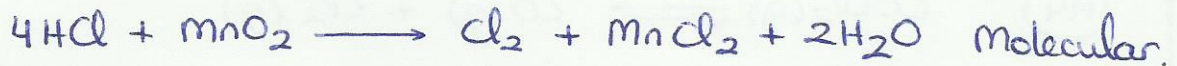
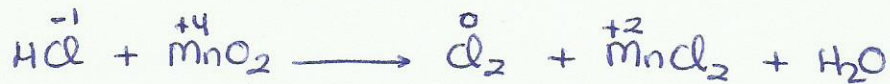


(A5)



b)  $V_{\text{HCl}}$

35% peso

$d = 1.17 \text{ g/cm}^3$

$0.5 \text{ g MnO}_2$

$$\text{moles MnO}_2 = \frac{\text{gr}}{\text{Pm}} = \frac{0.5}{87} = 5.75 \cdot 10^{-3} \text{ moles MnO}_2$$

$$\left| \begin{array}{l} 4 \text{ moles HCl} \text{ --- } 1 \text{ mol MnO}_2 \\ \quad \quad \quad \times \text{ --- } 5.75 \cdot 10^{-3} \end{array} \right.$$

$$x = 0.023 \text{ moles HCl}$$

$$\text{gr PUROS HCl} = \text{moles HCl} \cdot \text{Pm} = 0.023 \cdot 36.5 = 0.84 \text{ g PUROS HCl}$$

$$\% \text{ Riqueza} = \frac{\text{gr PUROS}}{\text{gr TOT}} \cdot 100$$

$$35 = \frac{0.84}{\text{g}_{\text{TOT}}} \cdot 100 \longrightarrow \text{g}_{\text{TOT}} \text{ HCl} = 2.4 \text{ g}$$

$$d = \frac{m}{\text{vol}} \longrightarrow 1.17 \text{ g/cm}^3 = \frac{2.4 \text{ g}}{\text{Vol (cm}^3)} \longrightarrow \boxed{\text{Vol} = 2.05 \text{ cm}^3}$$

↙  
Pasar litros.