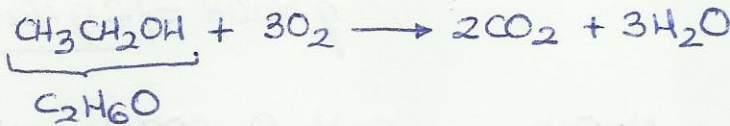


OPCIÓN B. PROBLEMAS

(B4)



$$215\text{g} \longrightarrow 75\text{KJ}$$

$$P_{m\text{C}_2\text{H}_6\text{O}} = 46\text{g/mol}$$

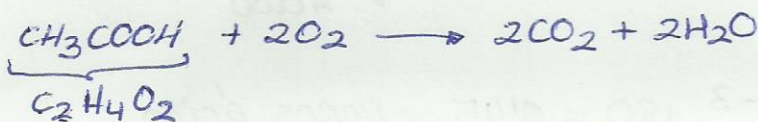
$$\text{moles} = \frac{215\text{g}}{46\text{g/mol}} = 0.054\text{ moles} \longrightarrow 75\text{KJ}$$

$$1\text{mol} \text{ --- } x$$

$$x = 1380\text{KJ.}$$

$$\text{Luego } \Delta H_{\text{comb etanol}}^\circ = \boxed{1380\text{KJ/mol}}$$

①



$$115\text{g} \longrightarrow 21\text{KJ}$$

$$P_{m\text{C}_2\text{H}_4\text{O}_2} = 60\text{g/mol}$$

$$\text{moles} = \frac{115\text{g}}{60\text{g/mol}} = 0.025\text{ moles} \text{ --- } 21\text{KJ}$$

$$1\text{mol} \text{ --- } x$$

$$x = 840\text{KJ.}$$

$$\text{Luego } \Delta H_{\text{comb ac}}^\circ = \boxed{840\text{KJ/mol}}$$

②



$$\Delta H_r^\circ = \Delta H_{\text{comb}_1}^\circ - \Delta H_{\text{comb}_2}^\circ$$